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therefore the reaction is exothermic or heat releasing 2  
answer a 1 explanation the entire reaction was multiplied by  
2 therefore  $\Delta h$  must also be halved 3 answer c explanation

option c is incorrect because 566 kj mol is the  $\Delta h$  per mole of  
o<sub>2</sub> and not for 2 moles

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both mass and enthalpy changes h<sub>2</sub> o s h<sub>2</sub> o l h 6 01 kj mol  
this is an endothermic process it requires 6 01 kj to melt one  
mole of ice h<sub>2</sub> o s the enthalpy value will change if the  
number of moles varies from the 1 1 reaction stoichiometry  
this is an exothermic process

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main concepts of thermochemistry such as the relationship  
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endothermic process heat capacity constant pressure  
calorimetry constant volume calorimetry the enthalpy the  
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web we will revisit this today and go on to see that if we use  
a special kind of thermochemical reaction called the  
standard enthalpy of formation we can calculate enthalpies  
of reactions without having to manipulate a series of

individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy  $\Delta U$  of a system in any process is equal to the heat  $q$  added to the system plus the work  $w$  done on the system by its surroundings  $e$

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