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Michal Rosen-Zvi

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cu in terms of these two isotopes answer 69 15 cu 63 and 30 85 cu 65 figure pageindex 5 isotopes of an element are atoms with the same atomic number but different mass numbers isotopes of an element therefore differ from each other only in the number of neutrons within the nucleus

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elements on the periodic table have at least two different natural isotopes many elements

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18 neutrons hence atomic mass 35
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 web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web jul 18 2018 answer link due to the number of neutrons isotopes differ with respect to the number of neutrons their nuclei contain where neutrons are massive particles of zero charge a nucleus may contain variable numbers of neutrons and this gives rise to the phenomenon of isotopes

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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