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than the OH^- alkaline solutions the H^+ of a basic solution is less
than $1 \times 10^{-7} \text{ M}$ they are the same as basic solutions

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web jan 10 2021 to identify the strongest base we can
determine their weakest conjugate acid the conjugate acids
of CH_3NH_2 and S^{2-} are CH_4NH_3^+ and HS^- respectively next we
consider that acidity increases with positive charge on the
molecule thus ruling out that S^{2-} is the weakest base

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web compounds formed from combinations of acids and
bases taste salty and are referred to in chemistry as salts
sweet compounds have characteristics of both acids and
bases in the same molecule we will explore the relationship
between molecular structure and acids bases and consider
water solutions of acids and bases

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loss or gain of a single hydrogen ion hydronium ion H_3O^+ a
water molecule that gains a hydrogen ion becomes this
which is positively charged what did lewis propose an acid
accepts a pair of electrons during a reaction while a base
donates a pair of electrons

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the ratio of the concentration of the dissociated form of an
acid to the undissociated form stronger acids have larger K_a
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problems pages 635 640 write balanced equations for

reactions between the following aluminum and sulfuric acid

$2\text{Al} + 3\text{H}_2\text{SO}_4(\text{aq}) \rightarrow 0\text{Al}_2(\text{SO}_4)_3(\text{aq}) + 3\text{H}_2(\text{g})$ calcium

carbonate and hydrobromic acid