

Casa Baldo 1915 Reseas

Raffaela Di Napoli

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web during this reaction the h^+ ions from the acid and the oh^-

ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web reacting two solutions eg acid and alkali place the
 polystyrene cup inside the glass beaker to make it more
 stable measure an appropriate volume of each liquid eg 25
 cm³ place one of the

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 measured enthalpy change of neutralization is a few kJ less
 exothermic than with strong acids and bases for example one
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 carbonates to produce salts neutralisation is the reaction
 between an acid and a base

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 when an acid is neutralised by an alkali the reaction is

exothermic eg $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = -57 \text{ kJ mol}^{-1}$
 eg $2 \text{H}_2\text{N} + \text{O}_3 \rightarrow \text{KOH(aq)} + \text{KNO}_3 \text{aq}$ for which $\Delta H = -57 \text{ kJ mol}^{-1}$

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 web the enthalpy change of neutralisation is defined as the
 enthalpy change when solutions of an acid and alkali react
 together to produce one mole of water for example $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(aq)}$ you can see that one mole of water is
 produced when one mole of hcl or one mole of naoh reacts
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