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web google classroom which of the
following statements accurately
describes the nernst equation shown
below $E_{\text{cell}} = E^\circ_{\text{cell}} - \frac{RT}{nF} \ln Q$
choose 1 answer choose 1 answer
choice a at equilibrium E_{cell}
calculates to a value of 0 and the
battery is

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web sep 27 2022 Cu^{2+} ions are lower
than H^+ ions in the electrochemical
series therefore selectively discharged
at the cathode Cu^{2+} ions have greater
tendency to
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web ans it suggests that Zn is more
reactive than hydrogen or that it is a
more powerful reducing agent zinc will
be oxidized to Zn^{2+} ions in a cell with a
zinc electrode and a

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web 10 $\log K = \frac{nE^\circ}{0.059}$ we can plug in the
value of E° on the left side of the
equation even though E° is normally
expressed as kJ mol^{-1} is expressed as J mol^{-1}
so we can convert E° or E to
match units let s plug in 300 000 J for E°
to match n divide 300 000 by
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web jul 16 2021 draw a well labelled
diagram of electrochemical cell formed
by half cells of metals p and q calculate
the voltage produced by the cell in ii
above when nitrate
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web sep 24 2020 ans 3 6 the cell in
which the following reaction occurs
 $2\text{Fe}^{3+} + 2\text{I}^- \rightarrow 2\text{Fe}^{2+} + \text{I}_2$ has $E^\circ_{\text{cell}} = 0.236 \text{ V}$ at 298 K calculate the standard

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web method 2 half reaction method
step 1 determine oxidation numbers for each atom step 2 use oxidation numbers to determine what is oxidized and what is reduced step 3

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web electrochemistry it is the study of the production of electricity from the energy released during spontaneous chemical reactions and the use of energy as non spontaneous

electrochemical reaction

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web reason when successive electrons are added to the orbitals in the same principal quantum level the shielding effect of inner core of electrons does not increase very much to

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web an electrochemical cell is any device that converts chemical energy into electrical energy or electrical energy into chemical energy there are

three components that make up an

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development of the word saucy its etymology traces back to the 1830s
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