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buffers typically consist of an acid base pair with the acid
and base differing by the presence or absence of a proton a
conjugate acid base pair for instance one of the buffers that
maintain the ph of human blood involves carbonic acid H_2CO_3
and its conjugate base the bicarbonate ion HCO_3^-

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an acid and a base react not to produce a salt and a solvent
but to form a new acid and a new base the concept of
neutralization is thus absent brønsted lowry acid base
behavior is formally independent of any solvent making it
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transfer of protons thus logic suggests a relation between
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strength of an acid or base is quantified in its ionization
constant K_a or K_b which represents the extent of the acid or
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acetate is a weak base $K_b = 5.6 \times 10^{-10}$ like acids
strong and weak bases are classified by the extent of their

ionization strong bases disassociate almost or entirely to
completion in aqueous solution

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stems from the recognition that some lewis acids and bases
seem to have a natural affinity for one another consider the
following some metals are commonly found in nature as salts
of chloride or as oxide ores while others are found in
combination with sulfur

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found in many every day items from fruit juices and soft
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oct 18 2023 most of the acids are solid it has a sour taste

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