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can be very useful if an acidic solution

of known concentration a standard

solution is added to a basic alkaline

solution of unknown concentration until

the solution is exactly neutralised i e

there is only salt and water it is

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in two forms an

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web the above equation works only for

neutralizations in which there is a 1 1

ratio between the acid and the base the

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is a quantitative analysis method used

to determine the concentration of an

acid or base by neutralizing the acid or

base with a known concentration

standard solution the concentration of a

solution can be determined by knowing

the acid and base dissociation constant

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when considering goals and procedures

however the most common types of

titration in quantitative chemical

analysis are redox titration and acid

base titration titrations can be

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titrations precipitation titrations

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titration methods which is a

quantitative method to analyze and

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concentration in the solution by

neutralizing them precisely with the

help of a standard solution which could

be of either base or acid of a known

concentration

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and base acid base titrations are

monitored by the change of ph as

titration progresses let us be clear

about some terminologies before we

get into the discussion of titration

curves titrant solution of a known

concentration which is added to

another solution whose concentration

has to be determined

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