

Matematicas Administrativas Unidad 1

S Nieto

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $2\text{Fe} + 3\text{CO} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{CO}_2$ oxidation $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$ reduction $\text{Fe} \rightarrow \text{Fe}_2\text{O}_3 + 2\text{e}^-$ $2\text{Fe} + 3\text{AgNO}_3 \rightarrow \text{Fe}_2(\text{NO}_3)_6 + 3\text{Ag}$ oxidation reduction $2\text{Fe} + 3\text{Cu} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{Cu}$ oxidation reduction $2\text{Fe} + 3\text{Cu} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{Cu}$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ $\text{Sr} \rightarrow \text{SrO} + 2\text{e}^-$ oxidized reducing agent $\text{O}_2 \rightarrow \text{O}^{2-}$ reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ $\text{Li} \rightarrow \text{Li}^+$ oxidized red ag $\text{S} \rightarrow \text{S}^{2-}$ reduced ox ag $3\text{Cs} + \text{Br}_2 \rightarrow 3\text{CsBr}$

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