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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $H_2C_2O_4(aq) + 2NaOH(aq) \rightarrow Na_2C_2O_4(aq) + 2H_2O(l)$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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concepts of stoichiometry and equivalence points we found
the concentration of an unknown substance by mixing 2m hcl
with the naoh of unknown concentration in order to
experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence
point considering the 1 1 stoichiometry of this acid base
reaction $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{NaC}_6\text{H}_4\text{COO} + \text{H}_2\text{O}$
the point of equivalence is the point of titration when the
number of moles of naoh added is equal to the number of
moles of khp

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