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**David Baud**

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$   
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 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ}$   
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2  $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{H}_2\text{O (l)} + \text{Na}^+ \text{(aq)}$   
3