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reactions oxidation cu s cu a 2 a q 2 e

reduction mno a 4 a a q 4 h a a q 3 e
mno a 2 s 2 h a 2 o l when the half
reactions listed above are combined
into a balanced chemical equation the
ratio of the coefficients cu s mno a 4 a a
q is

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redox reaction in basic solution no 2 aq
al s nh 3 g alo 2 aq 2 practice problem
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oxidation reaction in basic solution
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redox equation mno a 4 a aq so a 2 g
mn a 2 aq hso a 4 a aq what will be the
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way to balance redox reactions is by
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involves breaking an equation into its
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web solution 1 the two half reactions
balanced as if in acidic solution 2cn au
au cn 2 e 2e 2h o2 h2o2 2 make
electrons equal convert to basic
solution 4cn 2au 2au cn 2 2e multiplied
by a factor of 2 2e 2h2o o2 h2o2 2oh 3
the final answer

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balance any redox reaction is to
separate the reaction into half reactions
the substance being reduced will have
electrons as reactants and the oxidized
substance will have electrons as

products usually all reactions are
written as reduction reactions in half
reaction tables

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able to assign oxidation numbers well
oxidation the loss of electrons resulting
in a more positively charged species
reduction the gain of electrons
resulting in a

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web a redox equation can be balanced
using the following stepwise procedure
1 divide the equation into two half
reactions 2 balance each half reaction
for mass and charge 3 equalize the
number of electrons transferred in each
half reaction 4

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the oxidation and reduction of a redox
reaction in half reactions overall
scheme for the half reaction method
step 1 split reaction into half reactions
reduction and oxidation step 2 balance
the charge or oxidation number with

electrons step 3 balance c e o by adding c e h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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