

Cracking The Coding Interview 2022

RC Schank

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 matter in the previous section the reaction between
 hydrogen gas and oxygen gas to produce water in the
 gaseous phase was shown as a chemical equation h 2 g o 2 g
 h 2 o g at the molecular level the reaction would look
 something like this notice that there are two oxygen atoms
 on the

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 each type of element and polyatomic ion that are present in
 the reactants and the products of the reaction must be
 determined the first reactant in this equation contains
 magnesium mg and a hydroxide ion oh 1 which is a
 polyatomic anion

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 four simple steps to show the reactants and products write
 the imbalanced equation calculate the number of atoms of
 each element on each side of the reaction arrow to make the
 number of atoms of each element the same on both sides of
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 the carbon atoms first 5 use a coefficient to balance the
 single carbon atom add a coefficient to the single carbon
 atom on the right of the equation to balance it with the 3

carbon atoms on the left of the equation $c_3h_8o_2h_2o_3co_2$

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