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web equilibrium moles  $hscn 2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $hscn$  knowing the number of moles of  $fe^{3+}$  and  $hscn$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $mol\ fe^{3+} 1.70 \times 10^{-5} \text{ mol}$   $fe^{3+} 8.50 \times 10^{-4} \text{ m}$

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