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reaction are equal at equilibrium both  
reactants and products are present  
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at equilibrium some reaction mixtures  
contain mostly products and remain  
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equilibrium state is achieved when the  
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the pressure will usually the direction

that has fewer molecules  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$   
for every two molecules of  
ammonia made four molecules of  
reactant are used up this equilibrium  
shifts to the right with an increase in  
pressure

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decreases increases stays the same  
more products are formed in a chemical

reaction at equilibrium the rate of the forward reaction is equal to the rate of the reverse reaction if the rate of the forward reaction increases more products are formed

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