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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the

limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

experiment synthesis of 1 bromo 4

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion

bromobenzene will attack the nitronium ion by nucleophilic attack

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experiment was performed in order to

demonstrate the principles of nitration

through electrophilic aromatic

substitution it also examined the effects

of bromine on an aromatic ring ortho

para vs meso directing

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then actual is your actual findings yield

actual theoretical x 100 what are the

possible side reactions of nitration of

bromobenzene dinitration can occur at

ortho para or at both ortho positions

occurs in overheating meta product

happens very rarely becuae bromine is

an ortho para director

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the soap in the lab is not suitable for

this purpose 2 return to the lab and

apply mineral oil to the affected area

theory and explanation of the nitration

of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in

bromobenzene is an ortho para director

the para position is favored due to

steric effects in the reaction that takes

place in this lab a nitro group is

substituted onto bromobenzene at the

para position to form 1 bromo 4

nitrobenzene

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reacting sulfuric and nitric acid

together produces the nitronium ion

alone nitric acid is not a strong enough

electrophile to react with

bromobenzene sulfuric acid is added to

the nitric acid the reverse rxn can

overheat and boil easily includes a

claisen adaptor and

nitration of bromobenzene 6

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over 5 minutes slowly pipet this

bromobenzene into the icy cold

nitration mixture hno 3 h 2 so 4 in the

125 ml flask and stir continue to stir

the reaction mixture in the ice bath for

30 minutes do not heat in a new 250 ml

erlenmeyer

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on the right the arenium cation has a

nitro set with incredibly small 0 6

angstrom bond distances this will result

in an optimization failure as the atoms

are even close for the programme to

accomodate their electrons

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experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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