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web isotopes are atoms of the same element that differ in the amount of neutrons and atomic mass almost all elements on the periodic table have at least two different natural isotopes many elements

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number of protons isotopes are atoms that have the same number of protons but different numbers of neutrons thus atoms of $^{12}_6\text{C}$ and of $^{13}_6\text{C}$ are isotopes of each other

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therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation
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 web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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 web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic

table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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 web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope s mass number

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