

# Mike Bowditch Series In Order

**Judd E. Hollander**

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compounds  $\text{sc}_2$   $\text{n}_2\text{o}_4$   $\text{p}_4\text{o}_{10}$   $\text{pf}_5$   $\text{sc}_2$

sulfur dichloride  $\text{n}_2\text{o}_4$  dinitrogen

tetroxide the a of tetra is dropped

before the o of oxide  $\text{p}_4\text{o}_{10}$

tetraphosphorous decoxide the a of

deca is dropped before the o of oxide

$\text{pf}_5$  phosphorous pentafluoride

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h which immediately combines with a

water molecule to form the hydronium

ion h<sup>+</sup> the h<sup>+</sup> in is what defines the

acidic properties of a substance

because of the special classification

acids have a naming system different from

ionic or other molecular covalent

compounds

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molecules but when they are put into

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produced is always h<sup>+</sup> which immedi

ately combines with a water molecule

to form the hydronium ion h<sup>+</sup> the h<sup>+</sup> ion

is what defines the acidic properties

of a substance

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in model 2 that contain sulfur

phosphorus and nitrogen each pair has

one acid that ends in ic and another

that ends in ous these endings are

related to the name of the polyatomic

anion found in the acid ate or ite

key acid base part iii bronsted lowry

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each acid base reaction in model 2

describe the role of the bronsted lowry

base in the proton h<sup>+</sup> ion transfer that

occurs the base gains or accepts a

hydrogen ion proton as you saw in last

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model 1 have an oh<sup>-</sup> ion in their

chemical formulas

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acid chlorate h<sub>2</sub>so<sub>3</sub> answer sulfurous

acid sulfite h<sub>2</sub>so<sub>4</sub> answer sulfuric acid

sulfate h<sub>3</sub>po<sub>3</sub> answer phospho naming

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h which immediately combines with a

water molecule to form the hydronium

ion h<sup>+</sup> the h<sup>+</sup> ion is what defines the

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naming system different from ionic or

other molecular covalent compounds

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polyatomic anion ending is ite acid

name ending is 9 if the prefix hydro

were used to name a ternary acid what

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naming ternary acids 11 predict the

formula for chlorous acid

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acids have their own set of naming

rules cf sec 2 8 which you must also

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