

# Venus And Mars Book

**Ian Pickup**

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 equations 1  $\text{n}_2$   $\text{f}_2$   $\text{nf}_3$  2  $\text{c}_6\text{h}_{10}$   $\text{O}_2$   $\text{CO}_2$   $\text{H}_2\text{O}$  3  $\text{HBr}$   $\text{KHCO}_3$   $\text{H}_2\text{O}$   
 $\text{KBr}$   $\text{CO}_2$   $\text{GaBr}$   $\text{NaBr}$   $\text{SnO}$  4  $\text{Na}_2\text{SO}_3$   $\text{Ga}_2$   $\text{SO}_3$  3 5  $\text{NF}_3$   $\text{SnF}_2$   $\text{N}_2\text{O}_3$   
 using the equation from problem 2 above answer the  
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 balance the following equations 1  $\text{n}_2$   $\text{f}_2$   $\text{nf}_3$  2  $\text{c}_6\text{h}_{10}$   
 $\text{O}_2$   $\text{CO}_2$   $\text{H}_2\text{O}$  3 4 5 1  $\text{HBr}$  1  $\text{KHCO}_3$  1  $\text{H}_2\text{O}$  1  $\text{KBr}$  1  $\text{CO}_2$  2  
 $\text{GaBr}_3$  3  $\text{Na}_2\text{SO}_3$  1  $\text{Ga}_2$   $\text{SO}_3$  3 6  $\text{NaBr}$  3  $\text{SnO}$  2  $\text{NF}_3$  3  $\text{SnF}_2$  1  $\text{N}_2\text{O}_3$

using the following equation 2  $\text{NaOH}$   $\text{H}_2\text{SO}_4$  2  $\text{H}_2\text{O}$   $\text{Na}_2\text{SO}_4$   
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 calculator periodic table given the following reaction  $\text{Ce}$   $\text{Zn}$   
 $\text{CuCl}_2$   $\text{ZnCl}_2$   $\text{Cu}$   $\text{Zn}$   $\text{CuCl}_2$  2  $\text{ZnCl}_2$  2  $\text{Cu}$  how many moles of  $\text{Ce}$   
 $\text{ZnCl}_2$   $\text{ZnCl}_2$  2 will be produced from 23 0 text g 23 0 g of  $\text{Ce}$   
 $\text{Zn}$   $\text{Zn}$  assuming  $\text{Ce}$   $\text{CuCl}_2$   $\text{CuCl}_2$  2 is available in excess  
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 $\text{f}_2$  2  $\text{nf}_3$  br 2 2  $\text{c}_6\text{h}_{10}$  17  $\text{O}_2$  12  $\text{CO}_2$  10  $\text{H}_2\text{O}$  br 3 1  $\text{HBr}$  1  $\text{KHCO}_3$   
 1  $\text{H}_2\text{O}$  1  $\text{KBr}$  1  $\text{CO}_2$  br 4 2  $\text{GaBr}_3$  3  $\text{Na}_2\text{SO}_3$  1  $\text{Ga}_2$   $\text{SO}_3$  3 6  $\text{NaBr}$   
 br 5 3  $\text{SnO}$  2  $\text{NF}_3$  3  $\text{SnF}_2$  1  $\text{N}_2\text{O}_3$  br using the following  
 equation br 2  $\text{NaOH}$   $\text{H}_2\text{SO}_4$  2  $\text{H}_2\text{O}$   $\text{Na}_2\text{SO}_4$  br

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 equations 1 2 3 4 5 i  $\text{n}_2$  3  $\text{f}_2$  2  $\text{nf}_3$  2  $\text{CaH}_2$  40 17  $\text{O}_2$  12  $\text{CO}_2$  10  
 $\text{H}_2\text{O}$   $\text{HBr}$   $\text{KHCO}_3$   $\text{H}_2\text{O}$   $\text{KBr}$   $\text{CO}_2$   $\text{GaBr}$   $\text{Na}_2\text{SO}_3$   $\text{Ga}_2$   $\text{SO}_3$  3  $\text{NaBr}$   $\text{SnO}$   
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reaction and calculate how much  $\text{SiO}_2$  is required to produce 3.00 kg of  $\text{SiC}$  answer:  $\text{Ce SiO}_2 \rightarrow \text{SiC}$

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subscript 2 same thing goes with the  $\text{O}_3$  i am so lost

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