

# Fairyopolis Cicely Mary Barker

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the final temperature is closer to t<sub>1</sub>  
than to t<sub>2</sub> the final temperature is  
exactly halfway between t<sub>1</sub> and t<sub>2</sub> b  
the final temperature is exactly halfway  
between t<sub>1</sub> and t<sub>2</sub> the final  
temperature is closer to t<sub>2</sub> than to t<sub>1</sub> c  
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will go over the main types of questions  
on calorimetry including the heat  
capacity the heat of reaction finding the  
final temperature of a mixture constant  
pressure calorimetry and constant  
volume calorimetry a few important  
concepts and formulas you will need to  
solve these questions

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web calorimetry worksheet 1 if 0.315  
moles of hexane C<sub>6</sub>H<sub>14</sub> is combusted in  
a bomb calorimeter containing 5.65  
liters of water calculate the molar heat  
of combustion of hexane if the water  
temperature rises 55.4 °C the specific  
heat capacity of water is 4.184 J/g °C  
ms t h 5.650 grams H<sub>2</sub>O 4.184 J/g °C 55.4  
°C h 1310 kJ

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web 6 a 218 g sample of steam at 121 °C  
is cooled to ice at 14 °C find the change  
in heat content of the system 7 in going  
from ice at 34 °C to steam at 138 °C a  
sample of water absorbs 1.41 × 10<sup>5</sup> J  
find the mass of the sample answers 1 2  
68 × 10<sup>4</sup> J 2 6.23 × 10<sup>4</sup> J 3 1.71 × 10<sup>4</sup> J 4  
1.11 × 10<sup>5</sup> J 5 2.11 × 10<sup>5</sup> J 6

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