

Quotes From Domestic Violence Survivors

Camilla Rothe

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$\text{C}_2\text{H}_2\text{g} + 5\text{O}_2 \rightarrow 2\text{H}_2\text{g} + 2\text{H}_2\text{Og} + \Delta h$ 94
 5kJ $\text{H}_2\text{Og} + \text{H}_2\text{g} + \frac{1}{2}\text{O}_2\text{g} + \Delta h$ 71 2kJ $\text{C}_2\text{H}_6\text{g}$
 $7\text{O}_2\text{g} + 2\text{CO}_2\text{g} + 3\text{H}_2\text{Og} + \Delta h$ 283 kJg^{-1}
 $\text{C}_2\text{H}_6\text{g}$

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 worksheet calculate h for the reaction
 $\text{C}_2\text{H}_4\text{g} + \text{H}_2\text{g} \rightarrow \text{C}_2\text{H}_6\text{g}$ from the
 following data
 $\text{C}_2\text{H}_4\text{g} + 3\text{O}_2\text{g} \rightarrow 2\text{CO}_2\text{g} + 2\text{H}_2\text{Olg}$ 1411 kJmole^{-1}
 $\text{C}_2\text{H}_6\text{g} + 7/2\text{O}_2\text{g} \rightarrow 2\text{CO}_2\text{g} + 3\text{H}_2\text{Olg}$ 1560 kJmole^{-1}
 $\text{H}_2\text{g} + 1/2\text{O}_2\text{g} \rightarrow \text{H}_2\text{Olg}$ 285 kJmole^{-1}

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 the enthalpy change of an overall
 process is equal to the sum of the
 enthalpy changes of its individual steps
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 Δh for the target reaction
 $\text{C}_2\text{H}_4\text{g} + 3\text{O}_2\text{g} \rightarrow 2\text{CO}_2\text{g} + 2\text{H}_2\text{Olg}$ given the following
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 web libretexts states hess s law states
 that regardless of the multiple stages
 or steps of a reaction the total enthalpy
 change for the reaction is the sum of all

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 2019 which determines the enthalpy
 change value hess s law is written as
 $\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$ or $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

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 can be expressed as the sum of two or
 more steps the enthalpy change for the
 overall process is the sum of the Δh
 values for each step to use hess s law
 two principles must be understood one
 if an equation is reversed the sign of
 the Δh value is also reversed two if an
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 directions for each of the following
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 order to calculate the enthalpy change
 for the combustion of magnesium oxide
 $\text{Mg(s)} + 2\text{O}_2\text{(g)} \rightarrow \text{MgO(s)}$ we used

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web hess s law states that the total
 enthalpy change for the reaction will be
 the sum of all those changes no matter
 how many different steps or stages in
 the reaction there are cohen 2016 the
 equations for the reactions in the
 experiment done are as follows
 $\text{NaOH(s)} + \text{NaOH(aq)} \rightarrow \text{NaOH(aq)}$

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 experiment was designed using hess s
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 is the sum of two or more other
 reactions the Δh for the overall process
 must be the sum of the Δh values of the
 constituent reactions enthalpy change
 Δh is independent of the path that a
 reaction follows to move from reactants
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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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