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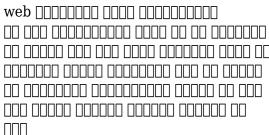
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product was determined to be ksp ca2 oh from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration 18 1 solubility product constant ksp chemistry libretexts - Feb 01 2023 web jul 12 2023 because the concentration of a pure solid such as ca 3 po 4 2 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore k ca2 3 po3 4 2 ca3 po4 2 ca3 po4 2 k ksp ca2 3 po3 4 2

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