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reaction needs to run to the right to achieve equilibrium b



what are the concentrations of all species at equilibrium  
check your final answers with a q calculation let x be the  
number of mol l of  $\text{H}_2$  or  $\text{I}_2$  that is lost to reach equilibrium  
 $\text{H}_2$  g  $\text{I}_2$  g  $2\text{HI}$  g

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of  $\text{NH}_3$  g became 0.540 m what is the value of the equilibrium  
constant for this reaction at the given temperature of 700 c  
 $\text{N}_2$  g  $3\text{H}_2$  g  $2\text{NH}_3$  i write the initial concentrations of  
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therefore we can describe the relationship between  $\Delta G^\circ$  and  
k p for gases as follows begin align 0  $\Delta G^\circ = -RT \ln K_p$  label 18  
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what does a dot represent 2 name two materials that the  
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that they were nonflexible 3 in model 1 the length of the  
arrows represents the average kinetic energy of the

molecules in that sample which gas variable p

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 $\text{g} + \text{I}_2$  g  $2\text{HI}$  g at 425oc an equilibrium mixture has the  
following concentrations  $\text{HI}$  1.01 x  $10^{-2}$  mol l  $\text{H}_2$  1.25 x  $10^{-3}$   
mol l  $\text{I}_2$  1.49 x  $10^{-3}$  mol l

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web equilibrium 1.864 m 1.364 m 0.636 m equilibrium 1.600  
m 1.000 m 0.400 m trial 5 a g b g c g trial 6 a g b g c g initial  
1.400 m 1.200 m 0.400 m initial 0.750 m 2.000 m 0.250 m  
change change equilibrium 1.388 m 1.188 m 0.412 m  
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