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web google classroom for the following redox equation  $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$  what will be the stoichiometric coefficient of  $\text{H}^+$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

10 4 balancing redox equations

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by the half reaction method by assuming a basic solution balancing redox reactions in basic solution chemteam - Sep 26 2022 web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + 2\text{e}^- + 2\text{H}_2\text{O} \rightarrow 2\text{HCN} + 2\text{OH}^-$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{e}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^-$  the final answer

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resulting in a

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1 divide the equation into two half reactions  
2 balance each half reaction for mass and charge  
3 equalize the number of electrons transferred in each half reaction  
4

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step 1 split reaction into half reactions reduction and oxidation  
step 2 balance the charge or oxidation number with electrons  
step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic

solution created by jay

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web bonus problem here is a balanced net ionic equation  
3br<sup>-</sup> 2 6oh<sup>-</sup> 5br<sup>-</sup> bro<sup>-</sup> 3 3h<sup>+</sup> 2 o the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons

are produced for every atom of nickel oxidized but where do those electrons go

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cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> fe<sub>2</sub> cr<sub>3</sub> fe<sub>3</sub> solution 1 balanced half reactions  
6e<sup>-</sup> 14h<sup>+</sup> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2cr<sub>3</sub> 7h<sub>2</sub>o  
fe<sub>2</sub> fe<sub>3</sub> e<sup>-</sup> 2 equalize the electrons  
6e<sup>-</sup> 14h<sup>+</sup> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2cr<sub>3</sub> 7h<sub>2</sub>o 6fe<sub>2</sub> 6fe<sub>3</sub> 6e<sup>-</sup>  
multiplied by a factor of 6

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web balance each half reaction in basic solution  
4 5 6 7 cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2 no so<sub>4</sub><sup>2-</sup> 2 mno<sub>2</sub> cr<sub>3</sub> no<sub>3</sub> so<sub>2</sub> mn<sub>2</sub>o<sub>3</sub> balance each redox reaction in acid solution using the half reaction method  
8 h<sub>2</sub>o<sub>2</sub> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2 9 teo<sub>3</sub>

2 n<sub>2</sub>o<sub>4</sub> o<sub>2</sub> cr<sub>3</sub> te 10 reo<sub>4</sub> 11 pbo<sub>2</sub> io i<sub>2</sub>  
pb<sub>2</sub> 12 as no<sub>3</sub> io<sub>3</sub> re io<sub>3</sub> h<sub>2</sub>aso<sub>4</sub>

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h<sub>2</sub> o<sub>2</sub> h and e in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no

choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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