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the objectives of the boiling point
determination lab were to determine
the boiling point of a mystery substance
provided to each student boiling point
is the point at which a liquid changes
physical states and becomes a vapor or
gas boiling point is important to
understand because it can be used to
help distinguish between different
substances

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the boiling point of an organic liquid is
the temperature at which its vapour
pressure equals the atmospheric
pressure over the liquid or it is the
temperature at which the vapour and
liquid phases are in equilibrium at
agiven pressure the boiling point is
considered as a criterion of purity of a
compound and is useful for
identification of

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given that the boiling points of ethanol
and methanol 78 c and 64 c
respectively are far off in a significant
way this seems like a valid conclusion

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are necessary in order to collect
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calculations are imperative to gathering
correct results

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melting point 1 a push the open end of
a capillary tube into the powdered
acetamide b move the powder to the
closed end of the capillary tube by
tapping it on the table repeat until the
the powdered acetamide occupies 1 2
mm of the capillary tube end c

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conclusion objectivesto correctly
measure boiling issues of various
liquids and identify an unknown liquid
based on its boiling point background
information objekte can be
distinguished into acid properties and
physical properties

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discussion and conclusion the

technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c
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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus
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melting point can be used as the identity and purity of an organic solid however it is not the same as the

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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