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to describe heat flow in terms of the change in a state
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spontaneously ΔG for the reaction must be less than zero in
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therefore the reaction is exothermic or heat releasing 2
answer a 1 explanation the entire reaction was multiplied by
2 therefore ΔH must also be halved 3 answer c explanation
option c is incorrect because 566 kJ/mol is the ΔH per mole of
 O_2 and not for 2 moles

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both mass and enthalpy changes $H_2O(s) \rightarrow H_2O(l)$ $\Delta H = 6.01 \text{ kJ/mol}$

this is an endothermic process it requires 6.01 kJ to melt one
mole of ice $H_2O(s)$ the enthalpy value will change if the
number of moles varies from the 1:1 reaction stoichiometry
this is an exothermic process

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more useful expressions is that the change in internal energy
 ΔU of a system in any process is equal to the heat q added to

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