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15.03 solution concentration

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molality $\frac{\text{moles solute}}{\text{kilograms solvent}}$
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2 molality $0.825 \text{ g} \times 0.141 \times 0.9579 \text{ g mol}^{-1} \times 0.00581158 \text{ mol} \times 0.00581158 \text{ mol} \times 0.4500 \text{ kg} \times 0.0129 \text{ m}$
3 mole fraction n_2 $\frac{\text{hpo}}{4}$

0.825 g $\times 0.141 \times 0.9579 \text{ g mol}^{-1} \times 0.00581158 \text{ mol}$
h 2 o 450.0 g $\times 0.18 \times 0.015 \text{ g mol}^{-1} \times 0.24 \times 0.97918401 \text{ mol}$

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 begin array l mass of a component $\frac{\text{mass of component in solution}}{\text{total mass of solution}} \times 100$ end array
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