

Stuttgart Akademie Der Bildenden Künste

Albert A Gayle

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the objectives of the boiling point determination lab were to determine the boiling point of a mystery substance provided to each student boiling point is the point at which a liquid changes physical states and becomes a vapor or gas boiling point is important to understand because it can be used to help distinguish between different substances

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed

end of the capillary tube by tapping it on the table repeat until the powdered acetamide occupies 1-2 mm of the capillary tube end c

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boiling point determination lab report conclusion
objectives to correctly measure boiling issues of various liquids and identify an unknown liquid based on its boiling point background information objects can be distinguished into acid properties and physical properties

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown compounds boiling points are affected by three key factors the first being intermolecular forces intermolecular forces affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to

identify substances or to check the purity of the compound it
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clean up solid and liquid spills immediately true g
 conclusions after conducting the boiling point determination
 lab it can be concluded that the unknown substance 193 is
 isopropyl alcohol on account of its observed boiling point of
 88 c

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conclusion after completing this experiment the melting
 point of three different unknown compounds was obtained
 using the mel temp apparatus compound a had a melting
 point range of 114 6 115 5 c thus revealing it to be
 acetanilide compound b had a melting point range of 122 6
 124 5 c thus

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nov 27 2019 essay sample the purpose of this experiment is
 to determine the boiling points of various organic compounds
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melting point can be used as the identity and purity of an
 organic solid however it is not the same as the boiling point
 the boiling point is a characteristic unique of pure liquids the
 boiling point is the temperature at which the vapor pressure
 of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in
 applied pressure so all boiling points should be reported with
 the measured pressure a compound s normal boiling point
 refers to its boiling point at a pressure of 760 text mm ce hg
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the purpose of the experiment is to understand what vapour
 pressure is and how it changes with temperature to
 understand the relationship between vapour pressure at a
 given temperature and boiling temperature at a given
 pressure to determine the vapour pressure of a pure liquid at
 various temperatures

exploring boiling points sharmin khatun chem 1141 studocu -
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exploring boiling points abstract in this experiment we are
 given a set of straight chain alkanes and a set of compounds
 to compare by following the trendline of the straight chain
 alkanes it could be concluded that as the molar mass of the
 straight chain alkanes increases their boiling points increase
 as well

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 conclusion read on should be based on observations made
 during the experiment

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laboratory partners samantha and mackenzie goals □
 correctly determine the boiling point of an unknown liquid □
 measure the refractive index of an unknown □ use the gas
 chromatogram to determine the mass percentage of a
 mixture and

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50 c so in order to convert an experimental measurement
 taken in calgary which is at higher altitude than sea level
 and so is at a lower pressure to that reported for sea level
 higher pressure one needs to add a correction factor since
 the boiling point at sea level is higher than that at higher
 altitudes

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introduction in chemistry one may encounter an unknown
 substance that need to be determined by physical properties
 such as boiling point melting point density solubility etc in
 this experiment i am to figure out the boiling point of a
 mystery substance in order to identify it chemical
 responsibility this experiment requires the handling of a
 bunsen burner boiling

6 2b step by step procedures for boiling point determination
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there are a variety of methods by which a sample s boiling
 point can be determined including distillation reflux and by
 using a thiele tube the most straightforward method uses a
 thiele tube and has the advantage of using less than 0 5ml 0
 5 ml of material

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 hizmetlerimizi sunmak müşterilerin hizmetlerimizi nasıl
 kullandığını anlayarak iyileştirmeler yapabilmek ve ilgi
 alanlarına göre özelleştirilmiş reklamlar da dahil olmak üzere
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mathematics is really about proving general statements via arguments usually called proofs

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